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Composition And Empirical Formulas
Part 2: Empirical Formulas Work each

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of the following problems. SHOW ALL WORK. 1. A compound is found to contain 63.52 % iron and 36.48 % sulfur. Find its empirical formula. 2. In the laboratory, a sample is found to contain 1.05 grams of nickel and 0.29 grams oxygen. Determine the empirical formula.

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~~7-21 % Composition and Empirical
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Nicotine is 74.1% carbon, 8.6% hydrogen, and 17.3% nitrogen by mass. Its molar mass is about 160 g./mol. 13. Epinephrine (adrenaline) is a hormone secreted into the

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bloodstream in times of danger and stress. It is 59.0% carbon, 7.1% hydrogen, 26.2% oxygen, and 7.7% nitrogen by mass.

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Unit 5 Worksheet 3: Empirical and
Molecular Formulas Complete these

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Formulas Answers
problems on notebook paper, or print this document and complete them below. Show all your work and be sure to include units and label your answer. Upload an image of your work to the assignment on Google Classroom.

~~05.07 WS3~~ — ~~Empirical and Molecular~~

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empirical formula. ? mol Fe = 63.52 g
Fe ! 1 mol Fe 55.8 g Fe = 1.14 mol
Fe? mol S = 36.48 g S ! 1 mol S 32.1
g S = 1.14 mol S Fe 1.14S 1.14 FeS 2.
In the laboratory, a sample is found to
contain 1.05 grams of nickel and 0.29
grams oxygen. Determine the

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Empirical formula. ? mol Ni = 1.05 g Ni
! 1 mol Ni 58.7 g Ni = 0.0179 mol Ni?
mol O = 0.29 g O ! 1 mol O 16.0 g O

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Name Date Pd Empirical Formulas
Worksheet, #1 Directions: Find the

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Empirical formula and name for each of the following. 1. A compound is 24.7% Calcium, 1.2% Hydrogen, 14.8% Carbon, and 59.3% Oxygen. Write the empirical formula and name the compound. 2. A compound is 21.20% Nitrogen, 6.06% Hydrogen, 24.30% Sulfur, and 48.45% Oxygen.

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~~SETH_SWEET_-_Empirical_Formula_
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Empirical Formula Worksheets. Some of the worksheets below are Empirical Formula Worksheets, several exam style questions like determine the empirical formula of a compound that

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contains 53.70% iron and 46.30% sulfur, ... solutions are provided at the end of each worksheet.

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About This Quiz & Worksheet. This quiz and corresponding worksheet will

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Formula Answers
Help you gauge your understanding of how to calculate percent composition and determine empirical formulas.

~~Quiz & Worksheet - How to Calculate Percent Composition ...~~

EMPIRICAL AND MOLECULAR
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Formulas Answers
Chromium is found to have the following % composition: 68.4 % Cr and 31.6 % O. Determine this compound's empirical formula. The percent composition of a compound was found to be 63.5 % silver, 8.2 % nitrogen, and 28.3 % oxygen. Determine the compound's empirical

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Empirical and Molecular Formula
Worksheet ANSWER KEY. Write the
empirical formula for the following
compounds. 1) C_6H_6 . C_3H_3 . 2)

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C_8H_{18} . C_4H_9 . 3) WO_2 . WO_2 . 4)
 $C_2H_6O_2$. CH_3O . 5) X_3Y_{13} . X_3Y .
6) A compound with an empirical
formula of $C_2O_2H_4$ and a molar mass
of 88 grams per mole. What is the
molecular formula of this compound?
 $C_4O_2H_8$. 7) A compound with an
empirical formula of C_4H_4O and a

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molar mass of 136 grams per mole.

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7. A compound with an empirical formula of CFBrO and a molar mass of 254.7 grams per mole. 8. A compound

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with an empirical formula of C_2H_8N
and a molar mass of 46 grams per
mole. Answer the following questions:

9. The percentage composition of
acetic acid is found to be 39.9% C,
6.7% H, and 53.4% O. Determine the
empirical formula of acetic ...

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Answers to Worksheet #8 Empirical

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Formulas To calculate empirical formulas, follow the steps outlined below: (assume percentages given in the problems are grams) Step 1: convert to moles Step 2: divide each by the lowest number of moles Step 3: (only if necessary) multiply all by the same factor in order to obtain whole

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~~Worksheet #8 Empirical Formulas H O
N O 4I~~

What is the empirical formula for a compound if a 2.50 g sample contains 0.900 g of calcium and 1.60 g of chlorine? ? mol Ca = 0.900 g Ca ! 1

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Mol Ca $40.1 \text{ g Ca} = 0.0224 \text{ mol Ca}$?
mol Cl = $1.60 \text{ g Cl} \div 1 \text{ mol Cl } 35.5 \text{ g Cl}$
= 0.0451 mol Cl Ca 0.0224 Cl 0.0451
CaCl₂ Calculate the % composition
for potassium sulfate. K_2SO_4 $2(39.1$
 $\text{g}) + 32.1 \text{ g} + 4(16.0 \text{ g} \dots$

~~Calculate the % composition for~~

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This activity worksheet examines both empirical formula and molecular formula through calculations and scaffolding steps. Students first calculate empirical formula using percentages and then progress into calculating molecular formula through examples.

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