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pH and pOH CONTINUED 1. 0.01 M HCl [H⁺] =

0.01M pH = $-\log(0.01) = 2$ 2. 0.0010 M NaOH

[OH⁻] = 0.0010M pOH = $-\log(0.010) = 3 \rightarrow$ pH =

11 3. 0.050 M NaOH

pH and pOH CONTINUED - Newbury Park High School

pH/pOH. pH and pOH are a measure of how acidic or how basic a solution is and for most solutions, is measured on a scale of 0 to 14. The lower the number, the more acidic the solution is, and the higher the number, the more basic the solution is. Neutral solutions (not acidic or basic) have a pH of 7.

pH / pOH - Chemistry Help

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